



Electric Batteries and Electrochemical Cells

For Prelims: Electrochemical Cells, [Electric vehicles](#), Voltaic Cells, [Lithium-ion Technology](#), [Nobel Prize in Chemistry for 2019](#), Applications of Electric Batteries.

For Mains: Electrochemical Cells, Evolutionary Trajectory of Batteries.

[Source: TH](#)

Why in News?

The [electric batteries](#) and **electrochemical cell advancements** have garnered significant attention for revolutionizing technology across sectors like [transport](#) and [energy](#), steering us toward a **sustainable future**.

What are Electric Batteries and Electrochemical Cells?

▪ Electric Batteries:

- An **Electric battery** is a device that **stores chemical energy** and converts it into electricity.
 - Batteries are made up of **one or more electrochemical cells** that are connected to **external inputs and outputs**.
- Electric batteries have transformed our world, **enabling the proliferation of motorization and wireless technology**.
- **Major Applications:**
 - **Portable Electronics:** Powering smartphones, laptops, tablets, and wearable devices.
 - **Transportation:** Driving [electric vehicles \(EVs\)](#) for both personal and public transportation, reducing reliance on fossil fuels.
 - **Renewable Energy Storage:** Storing energy **generated by solar panels and wind turbines for later use**.
 - **Electricity for Remote Areas:** Providing **electricity in remote or off-grid locations** where conventional power sources are unavailable or unreliable.

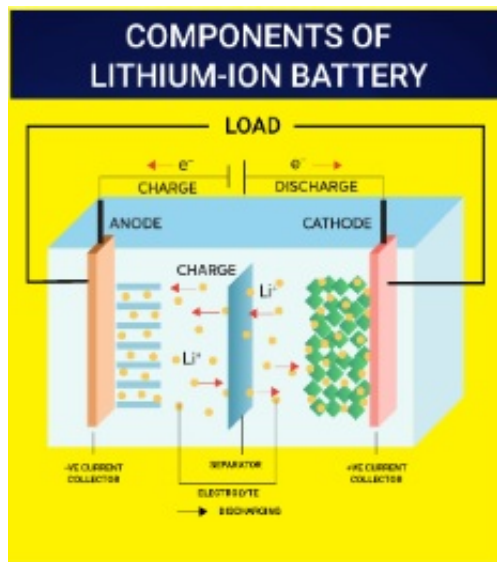
▪ Major Types of Batteries:

- **Solid-state battery:** It is a battery that uses solid electrodes and a solid electrolyte instead of a liquid or polymer gel electrolyte.
 - [Solid-state batteries](#) are used in a variety of devices, including: **pacemakers, radio frequency identifications (RFID) and wearable devices**.
- **Nickel-Cadmium battery (Ni-Cd):** They are used for **Cordless electronic appliances, drills, camcorders** and other small battery-operated devices requiring an even power discharge.
- **Alkaline Battery:** This is a type of primary battery that uses zinc and manganese dioxide as electrodes.
 - It is used for applications that require low cost and reliable power, such as **flashlights, toys, radios, and remote controls**.

- **Lithium-ion Battery:** The Li-ion battery's groundbreaking principles earned its developers the [Nobel Prize in Chemistry in 2019](#), underscoring its profound impact in the 20th and 21st centuries.
 - Li-ion batteries are versatile, **powering portable devices like phones and laptops** as well as fueling [electric vehicles](#) such as **cars and bikes**.
- **Electrochemical Cells:**
 - **Electrochemical cells** are devices that can **convert chemical energy into electrical energy, or vice versa**.
 - They can produce an electric current through chemical reactions, or they can use electrical energy to facilitate chemical reactions.
 - Electrochemical cells, like **voltaic or galvanic cells**, operate via **redox reactions** wherein **electrons are liberated during oxidation and utilized during reduction**.
 - A standard cell comprises two sections accommodating metal electrodes immersed in specific electrolytes.
 - The electrodes, namely the **anode and the cathode, conduct electricity**.
 - The **anode, where oxidation occurs**, and the **cathode, where reduction takes place**, form the fundamental components of the cell.
 - **Electrons flow from the negatively charged anode to the positively charged cathode** through an external circuit, providing power for a variety of uses.
 - Connecting these halves is a wire and a salt bridge, facilitating the movement of ions between them.
 - The **energy carried by electrons dictates the source voltage**, steering the electron flow within the circuit.
 - In ideal conditions, the **source voltage is equal to the terminal voltage**, ensuring an efficient power supply.
 - Advancements in cell design and materials, seen in **nickel-cadmium, zinc-copper, and modern lithium-ion cells**, showcase increased voltages and enhanced efficiency.
 - **Related Challenges:**
 - One of the well-known challenges affecting the efficiency of electrochemical cells is corrosion. For instance, in **environments with high humidity**, electrodes can gather water droplets.
 - If the atmospheric carbon dioxide levels are elevated, the **combination of water and gas leads to the formation of carbonic acid, causing corrosion** on the electrode surfaces.
 - Another issue arises from **galvanic corrosion**, where one of the electrodes within a cell **deteriorates faster in the electrolyte due to its higher reactivity**.
 - For instance, in a **carbon-zinc battery**, the zinc electrode erodes more rapidly during the battery's usage.

What is the Evolutionary Trajectory of Batteries?

- **Galvani's Experimentation (1780):**
 - Luigi Galvani's experiment **involving metals and frog legs** revealed a curious connection between **electrical energy and muscle movement**, laying the groundwork for future battery development.
- **Voltaic Pile (1800):**
 - Alessandro Volta's voltaic pile marked a significant step, **generating a steady current using metal plates and electrolytes**.
 - However, its functionality remained a mystery.
- **Faraday's Insights (Early 19th Century):**
 - Michael Faraday's groundbreaking work deciphered the mechanisms behind the cells, unveiling the roles of components like anode, cathode, and electrolyte.
- **Lithium-Ion Batteries:** This battery functions as both a voltaic and an electrolytic cell, capable of converting chemical energy into electrical energy and vice versa, enabling recharging.
 - In lithium-ion cells, **lithium metal oxide and graphite act as cathode and anode**, respectively, with a semisolid polymer gel electrolyte separating them.
 - The intercalation process enables charge and discharge phases.



Note: The Nobel Prize in Chemistry for 2019 was awarded to **John B. Goodenough, M. Stanley Whittingham, and Akira Yoshino** for their contributions to the development of the lithium-ion battery.

UPSC Civil Services Examination, Previous Year Question (PYQ)

Q. Which one of the following pairs of metals constitutes the lightest metal and the heaviest metal, respectively? (2008)

- (a) Lithium and mercury
- (b) Lithium and osmium
- (c) Aluminium and osmium
- (d) Aluminium and mercury

Ans: (b)

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